

X 1,5

19



Europäisches Patentamt  
European Patent Office  
Office européen des brevets

80110  
electrolyte  
no liquid

11 Publication number:

0 379 372

A1

list of patents  
w/ ionically conduc.  
powder in elec.

12

## EUROPEAN PATENT APPLICATION

21 Application number: 90300535.3

51 Int. Cl.<sup>5</sup>: H01M 6/18, H01M 8/10,  
H01M 10/36

22 Date of filing: 18.01.90

30 Priority: 18.01.89 US 298169

43 Date of publication of application:  
25.07.90 Bulletin 90/30

64 Designated Contracting States:  
DE ES FR GB IT

71 Applicant: MHB JOINT VENTURE  
3020 Newmark Drive  
Miamisburg Ohio 45342(US)

72 Inventor: Lee, Mei-Tsu  
15 Alley 3 Lane 770 Ming-Shung E. Road  
Taipei, Taiwan(CH)  
Inventor: Fauteux, Denis  
1016 Millerton Drive  
Centerville, Ohio 45459(US)

74 Representative: Deans, Michael John Percy et  
al  
Lloyd Wise, Tregear & CO. Norman House  
105-109 Strand  
London WC2R OAE(GB)

54 Composite solid electrolytes and electrochemical devices employing the same.

57 A composite solid electrolyte comprises a mixture of a solid ionically conductive powder and an ionically conductive polymeric material. The polymeric material provides a network containing the solid ionically conductive powder.

EP 0 379 372 A1

## COMPOSITE SOLID ELECTROLYTES AND ELECTROCHEMICAL DEVICES EMPLOYING THE SAME

The present invention relates to composite solid electrolytes, to their manufacture and to electrochemical devices employing the same.

Solid state electrochemical devices are the subject of intense investigation and development. They are described extensively in the patent literature. See, for example, U.S. Patents 4,303,748 to Armand; 4,589,197 to North; 4,547,440 to Hooper et al, and 4,228,226 to Christiansen. These cells are typically constructed of an alkali metal foil anode, an ionically conducting polymeric electrolyte solution containing an ionizable alkali metal salt, and a finely divided transition metal oxide as a cathode.

Solid electrolyte powders which are ionically conductive are disclosed in U.S. Patents 4,247,499 to Glugla et al.; 4,388,385 and 4,414,607 to Sekido et al.; 4,394,280 to Von Alpen et al.; 4,432,891 to Susman et al.; 4,539,276 to Harbach; and 4,557,985 to Voss.

The solid electrolyte in Glugla et al. 4,247,499 comprises a crystalline inorganic material embedded in a polymeric film. The preferred ionic conductive crystal is a beta-alumina crystal. The solid electrolyte in Sekido et al. 4,388,385 and 4,414,607 is based on a silver halide source. A particularly preferred electrolytic powder has the formula of  $\text{RbCu}_4\text{I}_{1.75}\text{Cl}_{3.25}$ . Matsushita Electric Industrial Company, assignee to the Sekido et al. patents, has developed a paper electrolyte wherein the electrolytic powder is mixed with a high polymer insulator material derived from styrene and butadiene and made into a cement. Von Alpen et al. 4,394,280 discloses a mixed crystal for use as an ion conducting solid electrolyte which is formed from components of  $\text{Na}_2\text{O}$ ,  $\text{ZrO}_2$ ,  $\text{P}_2\text{O}_5$  and  $\text{SiO}_2$ . Susman et al. 4,432,891, disclose a glass capable of ionic conduction. The glass is prepared from a non-metal glass former such as  $\text{GeS}_2$ ,  $\text{B}_2\text{S}_3$  and  $\text{SiS}_2$  in mixture with a glass modifier such as  $\text{Na}_2\text{S}$ . Voss 4,557,985 discloses a ceramic solid electrolyte. The electrolyte may take the form of a beta-alumina, a mixed crystal component, or a lithium nitride material.

Although the above-described solid electrolyte powders have been used in electrochemical devices, they suffer as a result of their powdery form. In practice, to produce a solid electrolyte material from the ionically conductive powders, the powders are typically compressed or compacted, by a mold, for example, to form a tightly adherent body. For some materials, particularly glass and ceramic materials, it is nearly impossible to form a unitary structure without the addition of materials such as binders. For other materials where a unitary structure may be produced, the resulting structure is extremely brittle and may be easily fractured as a result of its inflexible physical state.

It has been proposed to form an electrolyte material by forming an interpenetrating network from a curable material wherein the network functions to house an ionically conductive electrolyte. For example, Bauer et al. U.S. Patent 4,654,279 describes a cell in which the electrolyte is a two phase interpenetrating network of a mechanically supporting phase of a continuous network of a crosslinked polymer and an interpenetrating conducting liquid polymer phase comprising an alkali metal salt of a complexing liquid polymer which provides continuous paths of high conductivity throughout the matrix. In one embodiment, a liquid complex of a lithium salt and polyethylene oxide is supported by an epoxy, a polymethacrylate, or a polyacrylonitrile matrix. The network is formed by preparing a liquid solution of the metal salt, the salt-complexing liquid polymer, and the monomer for the crosslinked supporting phase in a polar solvent. The solvent is evaporated to form a dry layer of a mixture of the remaining materials. The dry layer is then cured.

Le Mehaute et al. U.S. Patent 4,556,614 discloses a solid electrolyte solution for an electrochemical cell in which a salt complexing polymer is mixed with a miscible and crosslinkable second polymer. The function of the second polymer is to maintain the complexing polymer in a more highly conductive amorphous state. This is accomplished by forming a solution of the two polymers and an ionizable salt in a solvent, evaporating the solvent, and crosslinking the second polymer. The second polymer is crosslinked by radiation.

Our European Patent Application 88310179.2 (Publication No. EP-A-0318161) describes a polymeric electrolyte wherein a radiation inert ionically conducting liquid having an ionizable alkali metal salt complexed therewith is maintained in a network of a cured photohardened polymer.

Although solid ionically conductive powders are known in the art, and the use of ionically conductive polymeric materials to form a matrix for housing an ionically conductive liquid or solid solution has been proposed, there has been no suggestion that an ionically conductive polymeric material be used to form a matrix for housing a solid ionically conductive powder for improving the mechanical properties of the electrolyte.

As will become clear from the detailed description which follows, solid electrolytes can be produced in accordance with the present invention having improved mechanical properties without a corresponding

reduction in ionic conductivity.

In accordance with a first aspect of the present invention, we provide a composite solid electrolyte characterised in comprising a mixture of a solid ionically conductive powder and an ionically conductive polymeric material wherein said polymeric material provides a matrix containing said solid ionically conductive powder.

The ionically conductive powder, which may be inorganic (e.g. glass or ceramic), may be mixed with an ionically conductive polymerizable or crosslinkable material which is subsequently cured to form the matrix.

In the most typical arrangement, the ionically conductive materials which form the matrix are low molecular weight compounds having at least one heteroatom capable of forming donor acceptor bonds with alkali metal ions, alkaline earth ions, silver ions, protons or proton-transporting ions. These materials may be polymeric in their natural form, or may be produced by curing a polymerizable or crosslinkable material which includes heteroatoms capable of forming donor acceptor bonds as described above. When the matrix forming material is produced by a curing procedure, the materials may be either radiation curable or thermally curable. When using a radiation curable ionically conductive material, exposure of the mixture produces an ionically conductive crosslinked or polymerized matrix which is interpenetrated by the solid ionically conductive powder. The support of the powder in the matrix provides an electrolyte material having improved mechanical strength and excellent ionic conductivity. For example, in comparison to systems using a nonconductive matrix material, the conductivity of the inventive electrolyte is improved by at least a factor of  $10^3$ .

In a second and alternative aspect of this invention, we provide a method for forming a composite solid electrolyte characterised in comprising the steps of: forming a mixture of solid ionically conductive powder and an ionically conductive polymeric material; and forming a matrix of said ionically conductive polymeric material such that said solid ionically conductive powder interpenetrates said matrix.

This method for producing a composite solid electrolyte is particularly advantageous because it can be used in the manufacture of the anode and cathode half elements as well as the manufacture of the electrochemical device produced from the anode and cathode half elements. Anode half elements can be prepared by coating the aforementioned mixture on a metal foil, for instance lithium metal on nickel or copper foil, and curing the coated foil, if necessary. The foil emerges with the composite solid electrolyte adhered to its surface. This not only provides intimate contact between the foil and the electrolyte but it also protects the underlying foil surface from damage during subsequent manufacturing operations in which it is assembled with the cathode element.

Similarly, the electrolyte material may be coated onto a cathode material and cured, if necessary. Alternatively, the cathode half element may be prepared by modifying the mixture to include the cathode material (e.g.,  $V_6O_{13}$ ) and a conductor (e.g., carbon black), coating the mixture on a metal foil support for the cathode half element, and curing the mixture, if necessary.

In forming completed electrochemical cells, the manufacture of the anode and the cathode half elements is merged into a single operation wherein the coated anode and cathode foil members are assembled. In one embodiment, after assembly, the cathode and anode members are cured together to form a completed cell. Cured anode and cathode half elements prepared as above can also be assembled using heat and pressure for lamination. Various other formats are also possible.

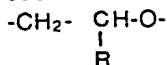
The invention provides, in a third alternative aspect thereof, a method for forming an electrochemical cell which comprises assembling an anode half element and a cathode half element with a solid composite electrolyte therebetween said solid composite electrolyte comprising a mixture of a solid ionically conductive powder and an ionically conductive polymeric material.

The term "solid ionically conductive powder" defines a material, which, in its solid form, is capable of ionically conducting alkali metal ions, alkaline earth ions, silver ions, protons or proton-transporting ions. The powders we employ are specifically to be distinguished from so called solid electrolytes which are actually solid solutions of an ionizable alkali salt complexed with a polymer having at least one hereto atom in its monomer pattern (e.g.  $LiClO_4$ /polyethylene oxide complex).

We have found that containment of the solid ionically conductive powder in the ionically conductive polymeric matrix produces an electrolyte which has excellent ionic conductivity and superior strength properties.

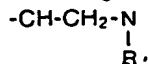
The matrix which is interpenetrated by the solid ionically conducting powder may be a polymeric material in its natural form or may take the form of a cured polymerizable or crosslinkable monomer.

General examples of useful polymers are described in U.S. Patent 4,303,748 to Armand and European Application 0 145 498 to Cook. These polymers have repeating units containing at least one heteroatom such as an oxygen or nitrogen atom. They can be represented as polymers having the repeating unit

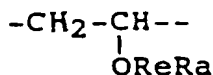


wherein R is hydrogen or a group Ra, -CH<sub>2</sub>ORa, -CH<sub>2</sub>OReRa, -CH<sub>2</sub>N(CH<sub>3</sub>)<sub>2</sub>, in which Ra is an alkyl group containing 1 to 16 carbon atoms and preferably 1 to 4 carbon atoms or a cycloalkyl group containing 5 to 8 carbon atoms, and Re is an ether group of formula -CH<sub>2</sub>-CH<sub>2</sub>Op- wherein p is a number from 1 to 100, preferably 1 or 2:

or having the repeating unit

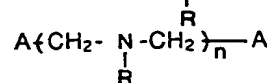
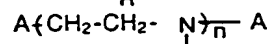
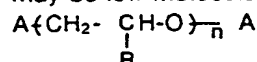


wherein R is Ra, or ReRa, as defined above; or having the repeating unit



wherein Re and Ra are as defined above. Copolymers of the above polymers may also be useful.

Ionically conductive curable materials useful in practice of the present invention include compounds having at least one and preferably a plurality of heteroatoms particularly oxygen and/or nitrogen atoms capable of forming donor acceptor bonds with alkali metal ions, alkaline earth ions, silver ions, protons or proton-transporting ions, and which are terminated by radiation polymerizable moieties. For example, they may be low molecular weight oligomers of the formulae:



where n is about 3 to 50 and R is hydrogen or a C1-C3 alkyl group which are terminated by ethylenically unsaturated moieties or glycidyl moieties represented by A. A particularly useful group of compounds is obtained by reacting a polyethylene glycol with acrylic or methacrylic acid. Polyethylene glycol diacrylate is a particularly preferred polymer. To provide additional structural integrity, triacrylate prepolymers may be added.

Preferably, the ionically conductive polymeric materials have a molecular weight of about 200 to 800. Still more preferably they are liquids at temperatures less than 30°C.

It may be desirable to include a radiation curable comonomer in the composition to reduce the glass transition temperature and improve conductivity of the polymer. Monoacrylate materials are particularly suitable for this purpose.

Also useful as comonomers in practice of the present invention are nonconductive radiation curable materials such as acrylated epoxies, e.g., Bisphenol A epoxy diacrylate, polyester acrylate, glycidyl ethers with acrylates, and a vinyl compound such as N-vinylpyrrolidone.

The solid ionically conductive powders which are maintained in the ionically conductive matrix are capable of ionically conducting alkali metal ions, earth alkali ions, protons or proton-transporting ions. These materials are inorganic in nature and often are glass or ceramic materials. Examples of solid ionically conductive powders which may be used in our electrolytes include RbAg<sub>4</sub>I<sub>5</sub>, RbCu<sub>16</sub>I<sub>17</sub>Cl<sub>13</sub>, Naβ-Al<sub>2</sub>O<sub>3</sub>, Ag<sub>5</sub>I<sub>2</sub>WO<sub>4</sub>, polycrystalline LiI, Na<sub>3</sub>Zr<sub>2</sub>Si<sub>2</sub>PO<sub>12</sub>, β-PbF<sub>2</sub>, LiI(Al<sub>2</sub>O<sub>3</sub>), and B<sub>2</sub>S<sub>3</sub>-Li<sub>2</sub>S-LiI. Other solid ionically conductive powders suitable for use in our electrolytes include those disclosed in U.S. Patent Nos. 4,247,499; 4,388,385; 4,394,280; 4,414,607; 4,432,891; 4,539,276; and 4,557,985. The solid ionically conductive powders typically have an ionic conductivity ranging between about 0.1 ohm<sup>-1</sup>cm<sup>-1</sup> and about 1.0 X 10<sup>5</sup> ohm<sup>-1</sup>cm<sup>-1</sup>.

The mixture of powder and polymer suitably contains about 50 to 99 weight percent solid ionically conductive powders and about 1 to about 50 weight percent ionically conductive polymeric material. The exact amount will vary with the nature of the polymeric material and the affinity of the polymeric material for the solid ionically conductive powders. As a general rule, at least 2 percent of the polymeric material must be added to demonstrate an appreciable improvement in mechanical properties.

To produce the composite solid electrolyte, the solid ionically conductive powders and the ionically conductive polymeric material are mixed together. In the case where the polymeric material is a radiation cured polymerizable or crosslinkable material, the mixture is passed through a source of actinic radiation. Similarly, if the polymeric material is a thermally cured polymerizable or crosslinkable material, the mixture

is heated to initiate polymerization. No materials other than the solid ionically conductive powders and the ionically conductive polymeric material need be present to form the composite solid electrolyte. However, as will be recognized by those skilled in the art, additives such as surfactants may be added to the mixture in minimal amounts.

5 The mixture may also include additional materials to provide additional improvements in properties such as ionic conductivity, mechanical strength, flexibility, and the like. An example of an additive material which is radiation curable is polyethylene oxide. To produce a composite electrolyte containing the additive polymer, the additive polymer is mixed into the mixture, the resultant material is extrusion coated onto a substrate, and the extruded material is passed through a source of actinic radiation to cure the mixture.

10 The term "actinic radiation" as used herein includes the entire electromagnetic spectrum and electron beam and gamma radiation. It is anticipated, however, based on availability of radiation sources and simplicity of equipment that electron beam and ultraviolet radiation will be used most often. Electron beam and gamma radiation are advantageous because they do not require the presence of a photoinitiator. When a photoinitiator is required, for example when using ultraviolet radiation, any conventional initiator may be used. When using electron beam, the beam potential must be sufficiently high to penetrate the electrode layer, the anode or cathode half element, or the cell itself depending upon which manufacturing technique is adopted. Voltages of 175 to 300 KV are generally useful. The beam dosage and the speed with which the element traverses the beam are adjusted to control the degree of crosslinking in an otherwise known manner.

20 Our method can be used to produce free standing solid thin electrolyte films or electrode half elements. To produce a free standing film, the mixture is poured into a mold or coated onto a surface having a release characteristic such as polytetrafluoroethylene. If the polymeric material is radiation curable, the mixture is cured by exposure to actinic radiation. The radiation curable mixture may also be coated on a metal foil such as aluminum foil, or can be cast in a container prior to curing. The film thickness can vary but films about 25 microns thick are useful in many applications. The obtained film can be assembled between cathode and anode half elements prepared by the processes disclosed herein or other processes and laminated under heat and pressure. A conductive adhesive may be used if necessary, although not required.

30 The composite solid thin film electrolyte produced has significantly improved mechanical properties when compared to prior solid electrolytes which are not maintained in a matrix of polymeric material. A problem we have found with pressed or formed solid ionically conductive powders (without polymeric materials) is their brittle nature which render them susceptible to cracks and fractures as a result of applied stress forces. By comparison, the present composite solid electrolytes are much stronger than the prior electrolytes and are much more flexible than the prior electrolytes. As a result, the present electrolytes are much more capable of withstanding stress forces without risking fracture. Further, by using an ionically conductive polymeric matrix, the ionic conductivity of the electrolyte is far superior to electrolytes comprising ionically conductive powder maintained in a non-conductive polymeric matrix. Increases in ionic conductivity by a factor of at least 1000 can be achieved by practical embodiments in accordance with the present invention as compared to electrolytes having ionically conductive powder maintained within a non-conductive matrix.

40 To manufacture an electrochemical cell, the electrolyte is placed between an anode material and a cathode material, and the materials are laminated together, typically under heat and pressure. The cell may optionally contain a current collector attached to the face of the cathode not contacting the electrolyte. If the polymeric material used to form the ionically conductive matrix is radiation curable, the electrochemical device can be assembled and then cured in situ. For example, in one arrangement, a lithium coated foil member can be coated with the radiation polymerizable electrolyte composition and over coated with the cathode coating composition described previously; or nickel foil can be coated with the cathode coating composition described previously and overcoated with the radiation polymerizable electrolyte composition. This structure can be cured by exposure to electron beam or another source of actinic radiation and the remaining electrically conductive foil member can be assembled with it. In another embodiment the latter foil member may be assembled with the structure prior to curing.

55 Anode half elements are obtained by coating a foil of the anode metal with the composite electrolyte. If the polymeric material is radiation curable, the coated foil is exposed to radiation. A typical foil is lithium foil or lithium coated foil such as nickel or copper foil having a layer of lithium deposited on its surface. Lithium is preferred because it is extremely electropositive and light in weight. The radiation curable composition may be coated onto the foil in any manner. The prepolymer selected which cures upon exposure to actinic radiation is extremely stable and does not chemically react with the lithium. Suitable techniques include rod coating, roll coating, blade coating, etc.

In a preferred embodiment the cathode coating composition contains the same polymer phase as binder as the polymeric material of the composite electrolyte. There is no phase separation between the cathode and the electrolyte and as a result, the interfacial resistance between these components of an electrochemical cell is significantly reduced.

Coating compositions for cathode half elements include particles of an insertion compound and an electrically conductive material along with the electrolyte composition, which functions as a dispersing medium for the cathode materials. A typical coating formulation for a cathode half element may contain about 50 to 80 parts of insertion compound, about 2 to 15 parts of an electrically conductive particle such as carbon black and about 15 to 50 parts of the ionically conductive electrolyte composition described above. The cathode half element is obtained by coating a foil member such as nickel foil in a thickness of about 10 to 100 microns with the aforesaid composition. If necessary, the cathode half element may be radiation or thermally cured. Alternatively, a polymer phase different than the electrolyte phase may function to bind the insertion compound and electrically conductive material to a substrate.

Insertion compounds and conductive particles useful in practical embodiments are well known in the art. Representative examples of insertion compounds are  $V_6O_{13}$ ,  $MoO_2$ ,  $MnO_2$  and  $TiS_2$ . Other examples can be found in the aforementioned references. A conductive particle is carbon black.

In accordance with a further embodiment, the composite cathodic particles described in U.S. Patent 4,576,883 to Hope can be dispersed in our electrolyte and coated on a metal foil member as described above.

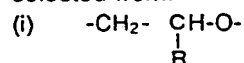
In preparing the coating compositions for the cathode half element, a small amount of a volatile solvent and a dispersing agent can be added to disperse the cathodic materials in the composition and produce a composition having good coating characteristics.

It is particularly envisioned that our electrolyte be used in an electrochemical cell. However, as will be appreciated in the art, the electrolyte may be used in other electrical components, such as capacitors.

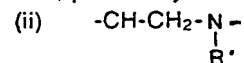
## Claims

1. A composite solid electrolyte characterised in comprising a mixture of a solid ionically conductive powder and an ionically conductive polymeric material wherein said polymeric material provides a matrix containing said solid ionically conductive powder.

2. An electrolyte according to Claim 1, further characterised in that material contains a repeating unit selected from:

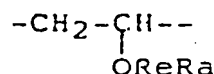


wherein R is hydrogen or a group  $R_a$ ,  $-CH_2OR_a$ ,  $-CH_2OR_eR_a$ , or  $-CH_2N(CH_3)_2$ , in which  $R_a$  is an alkyl group containing 1 to 16 carbon atoms and preferably 1 to 4 carbon atoms or a cycloalkyl group containing 5 to 8 carbon atoms, and  $R_e$  is an ether group of formula  $-CH_2-CH_2O_p-$  wherein p is a number from 1 to 100, preferably 1 or 2;



wherein  $R'$  is  $R_a$ , or  $ReR_a$ , as defined above; and

(iii)

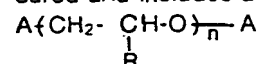


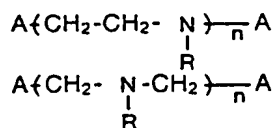
wherein  $R_e$  and  $R_a$  are as defined above, or said polymeric material comprises a copolymer containing two or more such repeating units.

3. An electrolyte according to Claim 1, further characterised in that said polymeric material comprises a cured polymerizable or crosslinkable material.

4. An electrolyte according to Claim 3, further characterised in that said polymeric material is thermally or radiation cured.

5. An electrolyte according to Claim 4, further characterised in that said polymeric material is radiation cured and includes a repeating unit of any of the following formulae:





- 5 where R is hydrogen or an alkyl group containing between 1 and 3 carbon atoms and n is an integer ranging between about 3 and about 50.
6. An electrolyte according to Claim 5, further characterised in that said polymeric material is ethylenically unsaturated.
7. An electrolyte according to any preceding claim, further characterised in that said ionically conduc-
- 10 tive powder comprises an inorganic powder, a glass powder or a ceramic powder.
8. An electrolyte according to Claim 7, further characterised in that said solid ionically conductive powder is selected from
- RbAg<sub>4</sub>I<sub>5</sub>, RbCu<sub>16</sub>I<sub>7</sub>Cl<sub>13</sub>, Naβ-alumina, Ag<sub>5</sub>I<sub>4</sub>WO<sub>4</sub>, polycrystalline LiI, Na<sub>3</sub>Zr<sub>2</sub>Si<sub>2</sub>PO<sub>12</sub>, β-PbF<sub>2</sub>, LiI(Al<sub>2</sub>O<sub>3</sub>) and B<sub>2</sub>S<sub>3</sub>-Li<sub>2</sub>S-LiI.
- 15 9. An electrolyte according to any preceding claim, further characterised in that said electrolyte further comprises an additional curable polymerizable or crosslinkable material.
10. An electrolyte according to Claim 9, further characterised in that said additional curable material comprises a nonconductive material.
11. An electrode half element characterised in comprising a metal foil having coated thereon an
- 20 electrolyte according to any preceding claim.
12. An element according to Claim 11, wherein said electrode is a cathode, further characterised in that said cathode additionally includes an insertion compound and electrically conductive particles.
13. An element according to Claim 11, wherein said electrode is an anode, further characterised in that said metal foil is a lithium foil or a lithium coated foil.
- 25 14. A method for forming a composite solid electrolyte characterised in comprising the steps of: forming a mixture of solid ionically conductive powder and an ionically conductive polymeric material; and forming a matrix of said ionically conductive polymeric material such that said solid ionically conductive powder interpenetrates said matrix.
15. A method according to Claim 14, further characterised in that said polymeric material is as defined
- 30 in any of Claims 2 to 6.
16. A method according to Claims 14 or 15, further characterised in that said conductive powder is as defined in any of Claims 7 or 8.
17. A solid state electrochemical cell including a composite solid electrolyte according to any of Claims 1 to 10.
- 35 18. A method for forming an electrochemical cell which comprises assembling an anode half element and a cathode half element with a solid composite electrolyte therebetween, said solid composite electrolyte comprising a mixture of a solid ionically conductive powder and an ionically conductive polymeric material.



DOCUMENTS CONSIDERED TO BE RELEVANT			EP 90300535.3
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int. Cl.)
A	<u>US - A - 4 752 544</u> (GREGORY) * Abstract *	1, 14	H 01 M 6/18 H 01 M 8/10 H 01 M 10/36
A	<u>US - A - 4 183 988</u> (FARRINGTON et al.) * Claims 1, 2 *	1, 7	
D, P, A	<u>EP - A1 - 0 318 161</u> (MHB) * Claims 1-4, 17, 19 *	1, 13	
D, A	<u>US - A - 4 654 279</u> (BAUER et al.) * Abstract *	1	
The present search report has been drawn up for all claims			TECHNICAL FIELDS SEARCHED (Int. Cl.)  H 01 M
Place of search VIENNA		Date of completion of the search 15-03-1990	Examiner LUX
<b>CATEGORY OF CITED DOCUMENTS</b> X : particularly relevant if taken alone Y : particularly relevant if combined with another document of the same category A : technological background O : non-written disclosure P : intermediate document T : theory or principle underlying the invention E : earlier patent document, but published on, or after the filing date D : document cited in the application L : document cited for other reasons & : member of the same patent family, corresponding document			